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# Mole Conversion Review Questions Chemistry Answer Key

**chemistry 12 some review of chem 11 - mole stuff** - chemistry 12 some review of chem 11 - mole stuff part 1 - changing moles to grams and grams to moles the two conversion factors to remember are: mm grams & 1 mole 1 mole mm grams where mm stands for the molar mass the molar mass is calculated by adding up atomic masses from underneath the symbol and the name on the periodic table. eg. the molar mass of  $\text{Na}_2\text{SO}_4$  is calculated as follows:  $2(23.0) + 3(16.0) + 4(32.1) = 233.0$  ... **the mole review - teachlearnchem** - the mole review \*\*\*\*all answers must include the proper units & sig figs.\*\*\*\* solve the following molar conversion & molarity problems: 1. how many grams would 8.1 10<sup>21</sup> molecules of sucrose (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>) weigh? 2. how many atoms are in 1.50 moles of iron? 3. how many atoms are in 127 grams of gold? 4. how many molecules are in 5.5 moles of  $\text{C}_2\text{H}_6$ ? 5. how many molecules are in 70.7 grams of  $\text{C}_4\text{H}_{10}$ ? 6. what is the mass in grams of  $5.25 \times 10^{24}$  molecules of  $\text{C}_2\text{H}_6$ ? ... **mole conversion problems - mr. bigler** - mole conversion problems note: some of these are the same compounds as in the "molar mass" worksheet, so you can use the formula weights from that worksheet as the starting point for your calculations. **extra review of mole conversions - mychemistryclass** - extra review of mole conversions directions: perform each conversion. show all work, show crossing units out, make sure answer has units. answers are in parenthesis. **chemistry\*11\*\* mole\*conversion\*practice\*test** - written: 1. how many atoms are in 1.50 moles of  $\text{C}_2\text{H}_6$ ? 2. what is the mass of a 250.0 ml sample of hydrogen sulfide ( $\text{H}_2\text{S}$ ) at STP? 3. at STP, how many liters of gas are produced from the reaction of 1 mole of  $\text{C}_2\text{H}_6$  with 1 mole of  $\text{O}_2$ ? ... **mole review practice problems** - mole practice problems 1. atomic mass vs. molar mass: ... write down the conversion factors that would be used to convert between the following units: molar mass (periodic table) avogadro's # ... **mole calculations review worksheet - answers on next page.** - chemistry honors name: \_\_\_\_\_ mole calculations review worksheet - answers on next page. 1. calculate the molar mass of each compound. **10-2 review and reinforcements - front page** - name date class 10—2 review and reinforcement mole conversions answer each of the following questions in the space provided. i. how would you calculate the number of moles present in a given mass of a substance? **chapter 2: moles, density, and concentration** - and 1 lb-mole of magnesium has a mass of 24.3 lb. we will use the designation of g mole for the si mole. molecular weight  $\equiv$  the sum of the atomic weights of the atoms that constitute a molecule of the compound molecular weight is mass per mole: molecular weight (mw) = mass/mole unit: kg/kmol, g/mol or lb m /lb-mole; molecular weight is the conversion factor that relates the mass and the ... **mole review practice problems - whitney high school** - 7. how much do  $3.01 \times 10^{23}$  atoms of he weigh? conv. factor(s) set-up 8. given 0.345 mole of water, how many grams would you have? **moles & stoichiometry cheat sheet** - moles & stoichiometry cheat sheet calculating molar mass 1. write out formula of compound 2. determine number of atoms of each element present **pre-ap chemistry chemical quantities review sheet** - page 1 of 4 pre-ap chemistry chemical quantities review sheet part 1: the mole and conversion - a conversion factor allows for the conversion between units. **mole to grams, grams to moles conversions worksheet** - 1 mole = molar mass (could be atomic mass from periodic table or molecular mass) 1 mole = 22.4 l of a gas at STP ( you do not need to worry about this yet ) each definition can be written as a set of two conversion factors.

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